Solutions.
Molarity = moles/L %(w/vol) = grams/100 mL, % (w/w) = g/100grams

If a solution is 5%(w/v) glucose, then there are 5 grams of glucose in 100 mL of the solution.

1. Convert ppm (1 part per million parts) to mass/volume

2. Convert ppb (1 part per billion parts) to mass/volume

3. What is the molarity of a solution that is made by diluting 10.0 mL of 0.2M sodium hydroxide to a total volume of 50 mL?

4. What is the molarity of a solution that is made by diluting 25.00 mL of 3.00M hydrochloric acid (HCl) to 500.0 mL?

5. Naming acids

Hydrochloric acid = HCl
Hydrobromic acid = HBr
Hydrofluoric acid = _____

Oxyacids are based on oxyanions:
Carbonic acid = H_2CO_3
Sulfuric acid = H_2SO_4 sulfurous acid = H_2SO_3
Nitric acid = HNO_3 nitrous acid = HNO_2

Hypochlorous acid = _____ Chlorous acid =_____

Chloric acid = _____. Perchloric acid = _____

6. Acid + Base → _____ + _____

What is the molarity of a sodium hydroxide solution, if 22.34 mL was needed to titrate 20.00 mL of a 0.1221 M solution of hydrochloric acid?