

CHEM 1100 Chapter 4 Titration worksheet. Dr. Stone

1. What volume (in mL) of 0.15M solution of glucose ($C_6H_{12}O_6$) has 3×10^{-3} moles of glucose?
a. 2×10^1 mL b. 2×10^{-2} mL c. 5×10^1 mL d. 5×10^4 mL e. None of these
2. How many moles of potassium phosphate are in 23.45 mL of 0.143 M solution of potassium phosphate?
a. 3.35 moles
b. 3.35×10^{-3} moles
c. 6.10×10^{-3} moles
d. 6.1 moles
e. None of these
3. It takes 32.56 mL of a sodium hydroxide solution to reach the end point of a titration of 25.00 mL of KHP. The concentration of KHP is 40.84g/L. What is the molarity of a sodium hydroxide solution?
4. It takes 22.12 mL of a sodium hydroxide solution to reach the end point of a titration of 20.00 mL of KHP. The concentration of KHP is 20.11g/L. What is the molarity of a sodium hydroxide solution?
5. It takes 22.12 mL of a sodium hydroxide solution to reach the end point of a titration of 20.00 mL of KHP. The concentration of KHP is 20.11g/L. What is the molarity of a sodium hydroxide solution?