CHEM 1100 Worksheet for Chapter 2: Naming compounds Dr. Stone

1. Name these polyatomic ions

   CH_3COO^-_____

   CO_3^{-2}_____

   HCO_3^-_____

   CN^{-1}_____

2. Name these polyatomic ions:

   ClO^{-1}_____

   ClO_2^{-1}_____

   ClO_3^{-1}_____

   ClO_4^{-1}_____

3. Name these polyatomic ions:

   $Cr_2O_7^{-2}$ _____ CrO_4^{-2}_____

   MnO_4^{-1} _____ N_3^{-1}_____

4. What is the correct formula/name for the following compounds:

   Sodium dichromate_____ Na_2O_2_____

   Magnesium permanganate_____ CuO _____

   Potassium azide_____ LiF _____

   Iron(III) chromate_____ NH_4Cl_____

   $Ca(NO_3)_2$_____

5. Name these compounds:

   $Ca(NO_3)_2$_____

   Na_2O_2_____

   Cu_2O_____

   LiOH_____

   $(NH_4)_2S$_____

6. What is the correct name for $CaCO_3$?_____

7. What numbers go with these prefixes?
We only use these when there are two non metals.
Di_____
Tri_____
Tetra_____
Penta_____
Hexa _____
Hepta_____
Octa_____
Nona_____
Deca_____

8. Name these compounds (there are 2 non metals)
PCl_5_____
NO_2_____
N_2O_____
N_2O_5_____
SO_3_____

9. What is the correct name for $Mg(NO_3)_2$ _____

10. What is the correct formula for lithium phosphate?_____

11. What is the correct name for $Co_3(PO_4)_2$?_____

12. Write the correct symbols for
proton _____

neutron_____

electron _____

13. What does the number under each element mean?

14. There are two isotopes of copper.
^{63}Cu has a mass of 62.9296 amu
^{65}Cu has a mass of 64.9278 amu
The weighted average mass of both isotopes of Copper is 63.546 amu
What is the percent of each isotope?