

1. 1.37 M
2. 1.04 M
3. 6.7×10^{-1} L
4. 3.3×10^{-1} L
5. 0.55M
6. 2.5×10^{-1} mL
7. 1.1×10^{-1} M
8. 5.00g HCl
9. $C_7H_6O_3$
10. 7.2×10^{-3} moles of acid
11. Should be: use the information (1 gram) and the moles you obtained in **question 10** (7.2×10^{-3} moles of acid) to determine the Molar mass (g/mole)
= 1.38×10^2 g/mole
12. $3 Ca^{+2}_{(aq)} + 2 PO_4^{-3}_{(aq)} \rightarrow Ca_3(PO_4)_2 (s)$
(Two soluble ions form an insoluble precipitate.)